

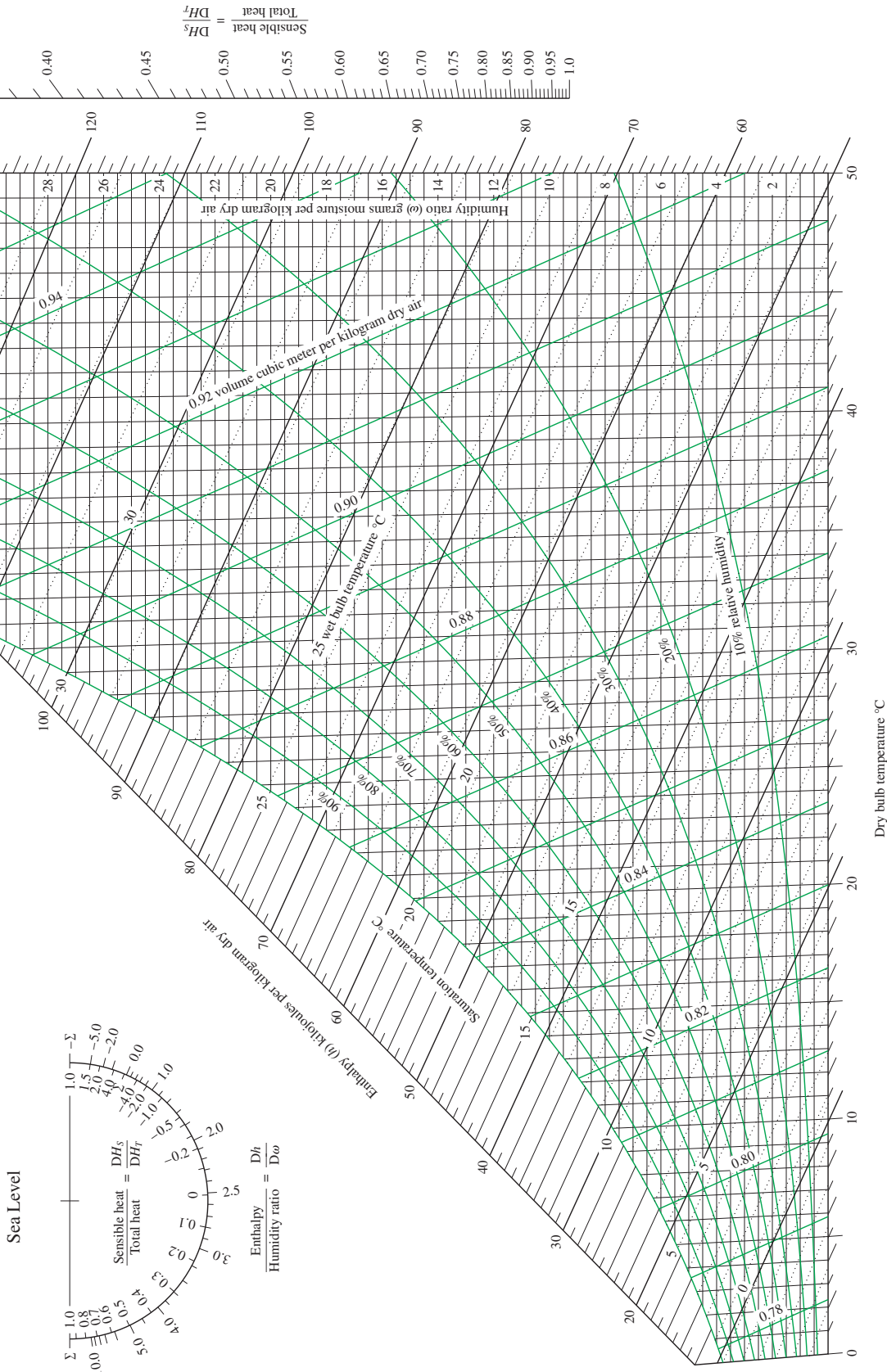
**FIGURE A-30**  
Generalized entropy departure chart.

Source of Data: Redrawn from Gordon van Wylen and Richard Sontag, Fundamentals of Classical Thermodynamics, (SI version), 2d ed., Wiley, New York, 1976.



**ASHRAE Psychrometric Chart No. 1**  
 Normal Temperature  
 Barometric Pressure: 101.325 kPa

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Prepared by Center for Applied Thermodynamic Studies, University of Idaho.

**FIGURE A-31**

Psychrometric chart at 1 atm total pressure.

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**TABLE A-32**

One-dimensional isentropic compressible-flow functions for an ideal gas with  $k = 1.4$

Ma	Ma*	A/A*	P/P <sub>0</sub>	ρ/ρ <sub>0</sub>	T/T <sub>0</sub>
0	0	∞	1.0000	1.0000	1.0000
0.1	0.1094	5.8218	0.9930	0.9950	0.9980
0.2	0.2182	2.9635	0.9725	0.9803	0.9921
0.3	0.3257	2.0351	0.9395	0.9564	0.9823
0.4	0.4313	1.5901	0.8956	0.9243	0.9690
0.5	0.5345	1.3398	0.8430	0.8852	0.9524
0.6	0.6348	1.1882	0.7840	0.8405	0.9328
0.7	0.7318	1.0944	0.7209	0.7916	0.9107
0.8	0.8251	1.0382	0.6560	0.7400	0.8865
0.9	0.9146	1.0089	0.5913	0.6870	0.8606
1.0	1.0000	1.0000	0.5283	0.6339	0.8333
1.2	1.1583	1.0304	0.4124	0.5311	0.7764
1.4	1.2999	1.1149	0.3142	0.4374	0.7184
1.6	1.4254	1.2502	0.2353	0.3557	0.6614
1.8	1.5360	1.4390	0.1740	0.2868	0.6068
2.0	1.6330	1.6875	0.1278	0.2300	0.5556
2.2	1.7179	2.0050	0.0935	0.1841	0.5081
2.4	1.7922	2.4031	0.0684	0.1472	0.4647
2.6	1.8571	2.8960	0.0501	0.1179	0.4252
2.8	1.9140	3.5001	0.0368	0.0946	0.3894
3.0	1.9640	4.2346	0.0272	0.0760	0.3571
5.0	2.2361	25.000	0.0019	0.0113	0.1667
∞	2.2495	∞	0	0	0

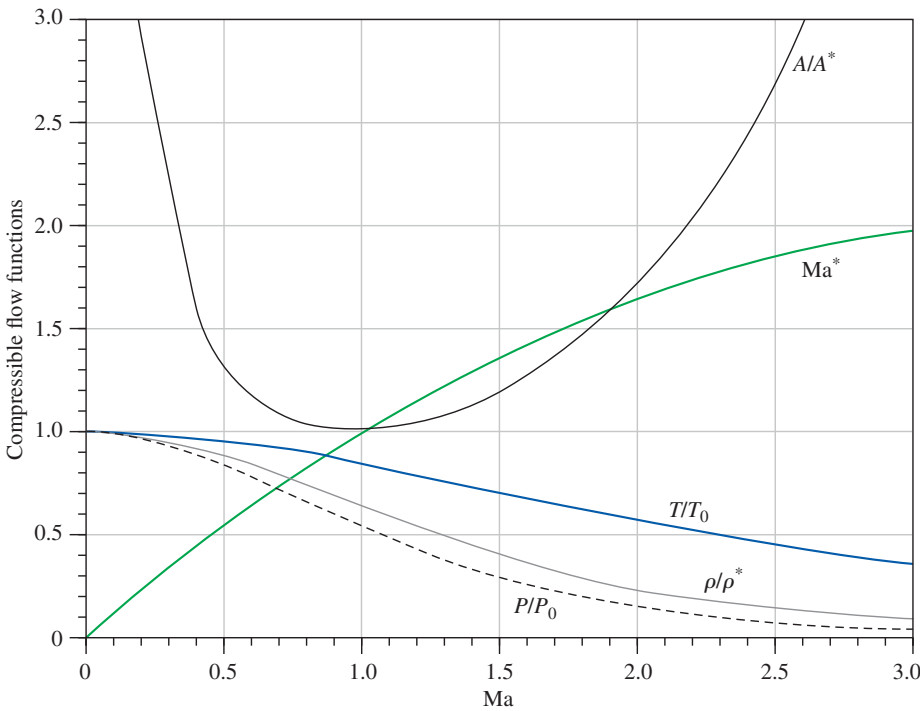
$$Ma^* = Ma \sqrt{\frac{k+1}{2+(k-1)Ma^2}}$$

$$\frac{A}{A^*} = \frac{1}{Ma} \left( \frac{2}{k+1} \right) \left( 1 + \frac{k-1}{2} Ma^2 \right)^{0.5(k+1)/(k-1)}$$

$$\frac{P}{P_0} = \left( 1 + \frac{k-1}{2} Ma^2 \right)^{-k/(k-1)}$$

$$\frac{\rho}{\rho_0} = \left( 1 + \frac{k-1}{2} Ma^2 \right)^{-1/(k-1)}$$

$$\frac{T}{T_0} = \left( 1 + \frac{k-1}{2} Ma^2 \right)^{-1}$$



$$T_{01} = T_{02}$$

$$Ma_2 = \sqrt{\frac{(k-1)Ma_1^2 + 2}{2kMa_1^2 - k + 1}}$$

$$\frac{P_2}{P_1} = \frac{1 + kMa_1^2}{1 + kMa_2^2} = \frac{2kMa_1^2 - k + 1}{k + 1}$$

$$\frac{\rho_2}{\rho_1} = \frac{P_2/P_1}{T_2/T_1} = \frac{(k+1)Ma_1^2}{2 + (k-1)Ma_1^2} = \frac{V_1}{V_2}$$

$$\frac{T_2}{T_1} = \frac{2 + Ma_1^2(k-1)}{2 + Ma_2^2(k-1)}$$

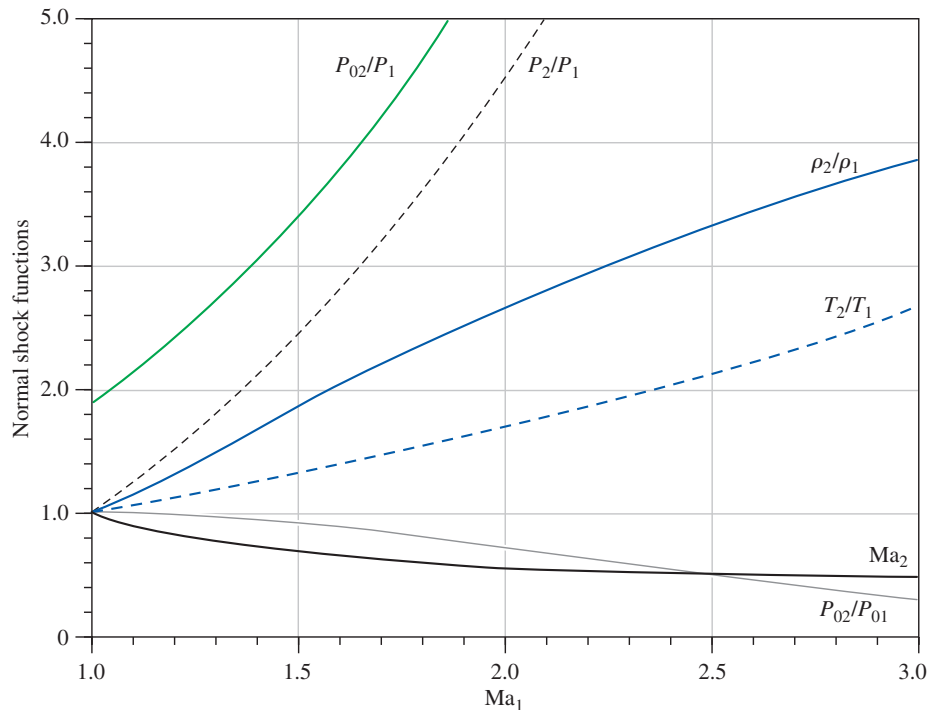
$$\frac{P_{02}}{P_{01}} = \frac{Ma_1 \left[ 1 + Ma_2^2(k-1)/2 \right]^{(k+1)/(2(k-1))}}{Ma_2 \left[ 1 + Ma_1^2(k-1)/2 \right]}$$

$$\frac{P_{02}}{P_1} = \frac{(1 + kMa_1^2) \left[ 1 + Ma_2^2(k-1)/2 \right]^{k/(k-1)}}{1 + kMa_2^2}$$

**TABLE A-33**

One-dimensional normal-shock functions for an ideal gas with  $k = 1.4$

Ma <sub>1</sub>	Ma <sub>2</sub>	P <sub>2</sub> /P <sub>1</sub>	ρ <sub>2</sub> /ρ <sub>1</sub>	T <sub>2</sub> /T <sub>1</sub>	P <sub>02</sub> /P <sub>01</sub>	P <sub>02</sub> /P <sub>1</sub>
1.0	1.0000	1.0000	1.0000	1.0000	1.0000	1.8929
1.1	0.9118	1.2450	1.1691	1.0649	0.9989	2.1328
1.2	0.8422	1.5133	1.3416	1.1280	0.9928	2.4075
1.3	0.7860	1.8050	1.5157	1.1909	0.9794	2.7136
1.4	0.7397	2.1200	1.6897	1.2547	0.9582	3.0492
1.5	0.7011	2.4583	1.8621	1.3202	0.9298	3.4133
1.6	0.6684	2.8200	2.0317	1.3880	0.8952	3.8050
1.7	0.6405	3.2050	2.1977	1.4583	0.8557	4.2238
1.8	0.6165	3.6133	2.3592	1.5316	0.8127	4.6695
1.9	0.5956	4.0450	2.5157	1.6079	0.7674	5.1418
2.0	0.5774	4.5000	2.6667	1.6875	0.7209	5.6404
2.1	0.5613	4.9783	2.8119	1.7705	0.6742	6.1654
2.2	0.5471	5.4800	2.9512	1.8569	0.6281	6.7165
2.3	0.5344	6.0050	3.0845	1.9468	0.5833	7.2937
2.4	0.5231	6.5533	3.2119	2.0403	0.5401	7.8969
2.5	0.5130	7.1250	3.3333	2.1375	0.4990	8.5261
2.6	0.5039	7.7200	3.4490	2.2383	0.4601	9.1813
2.7	0.4956	8.3383	3.5590	2.3429	0.4236	9.8624
2.8	0.4882	8.9800	3.6636	2.4512	0.3895	10.5694
2.9	0.4814	9.6450	3.7629	2.5632	0.3577	11.3022
3.0	0.4752	10.3333	3.8571	2.6790	0.3283	12.0610
4.0	0.4350	18.5000	4.5714	4.0469	0.1388	21.0681
5.0	0.4152	29.0000	5.0000	5.8000	0.0617	32.6335
∞	0.3780	∞	6.0000	∞	0	∞





**TABLE A-34**

Rayleigh flow functions for an ideal gas with  $k = 1.4$

Ma	$T_0/T_0^*$	$P_0/P_0^*$	$T/T^*$	$P/P^*$	$V/V^*$
0.0	0.0000	1.2679	0.0000	2.4000	0.0000
0.1	0.0468	1.2591	0.0560	2.3669	0.0237
0.2	0.1736	1.2346	0.2066	2.2727	0.0909
0.3	0.3469	1.1985	0.4089	2.1314	0.1918
0.4	0.5290	1.1566	0.6151	1.9608	0.3137
0.5	0.6914	1.1141	0.7901	1.7778	0.4444
0.6	0.8189	1.0753	0.9167	1.5957	0.5745
0.7	0.9085	1.0431	0.9929	1.4235	0.6975
0.8	0.9639	1.0193	1.0255	1.2658	0.8101
0.9	0.9921	1.0049	1.0245	1.1246	0.9110
1.0	1.0000	1.0000	1.0000	1.0000	1.0000
1.2	0.9787	1.0194	0.9118	0.7958	1.1459
1.4	0.9343	1.0777	0.8054	0.6410	1.2564
1.6	0.8842	1.1756	0.7017	0.5236	1.3403
1.8	0.8363	1.3159	0.6089	0.4335	1.4046
2.0	0.7934	1.5031	0.5289	0.3636	1.4545
2.2	0.7561	1.7434	0.4611	0.3086	1.4938
2.4	0.7242	2.0451	0.4038	0.2648	1.5252
2.6	0.6970	2.4177	0.3556	0.2294	1.5505
2.8	0.6738	2.8731	0.3149	0.2004	1.5711
3.0	0.6540	3.4245	0.2803	0.1765	1.5882

$$\frac{T_0}{T_0^*} = \frac{(k + 1)Ma^2[2 + (k - 1)Ma^2]}{(1 + kMa^2)^2}$$

$$\frac{P_0}{P_0^*} = \frac{k + 1}{1 + kMa^2} \left( \frac{2 + (k - 1)Ma^2}{k + 1} \right)^{k/(k-1)}$$

$$\frac{T}{T^*} = \left( \frac{Ma(1 + k)}{1 + kMa^2} \right)^2$$

$$\frac{P}{P^*} = \frac{1 + k}{1 + kMa^2}$$

$$\frac{V}{V^*} = \frac{\rho^*}{\rho} = \frac{(1 + k)Ma^2}{1 + kMa^2}$$

